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## IN THE CLAIMS

Please see attachment.

## RESPONSE

Applicants respectfully disagree with the statement that support has not been provided for the amended structure of formula I. However, in furtherance of allowing the application, applicants will cancel the structure amendment in the specification and in the claims.

In fact, the original figure as shown on page 3 was intended to show a polymeric form of the original monomer. However, the brackets were drawn incorrectly. This was obviously a clerical error, such that the brackets were merely drawn in the wrong place .

An examination of the pathways as taught by the disclosure, and an understanding of the chemical nature of a polymer that would follow from the chemical steps taught in the disclosure, would naturally lead one to the conclusion that in fact no new matter is being submitted, and that only a typographical correction is being made. This typographical error was merely a CLERICAL ERROR.

The chemistry of the reactions and the steps leading up to the reaction are supportive of this position. If one carefully follows the steps of the reaction as claimed and as taught in the specification, the

reaction would not lead to a polymer resin having a lengthy polymer up to 500 units, with each end of the polymer being "capped" by an aromatic structure or molecule. Indeed, it would be a virtual chemical impossibility to have reactions taking place in a large reactor vat wherein a lengthy polymer is formed with the condition that a single aromatic molecule or compound is attached at each end of said polymer.

The formation as claimed wherein an aromatic structure forms a bond with each nitrogen atom of the ring is less stereochemically and chemcially objectionable, and falls within the confines of polymer chemistry. Indeed, given the level of reactivity between a nitrogen atom and an aromatic compound, it become evident that in a situation of competitive chemical formation during a reaction, the aromatic structure would form bond with the nitrogen atom.

In the typing of the and preparation of the original application, it is obvious that the clerical error was the result of trying to showing the continuous nitrogen-aromatic bonding of the polymer, and not an effort to show an aromatic-aromatic bonding.

In furtherance of obtaining an allowance, most of the Office Action's issues have been addressed however, with respect to claim 34, the claimed temperature is found on page 7, line 5.

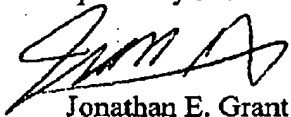
With respect to step A(iv) the reaction solution is cooled to 20 C. This is more of a translational issue. Support is found by the fact that the solution has been at a temperature of 58-63 degrees, C. as indicated on page 5, stage A, step b.

Claim 27 section A(iv) has been amended. The phrase "drinking water" only meant that the water did not have to be distilled. This has been changed, and the word "the" has also been removed. Also, the term "bromo" has been substituted with the term "halogen."

Applicants do not believe that any revoke any proposed amendments in the specification that the examiner considers new matter, if said revocation will lead to the allowance of the application..

The application is now in condition for allowance. Please call or fax me at (301) 603-9071 if you have any questions or comments.

Respectfully submitted,



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IN THE CLAIMS

Claims 1-36 (canceled)

-37 (New) . A process for obtaining polyglycolyl urea resin from aromatic diglycinates for insulating electric conductor, in the absence of HCN polluting residues, comprising the following steps:

A) preparing a methyl diglycinate:

- (i) reacting a mixture of methylhaloester and methylenedianiline in the presence of  $C_1-C_4$  aliphatic solvent under reflux conditions at atmospheric pressure at a solvent reflux temperature of  $58-63^{\circ}\text{C}$ , wherein said methylhaloester is selected from the group consisting of methylbromopropionate and methylchloropropionate;
- (ii) adding triethylamine, at a rate of 0.178 l/hr. per Kg of product;
- (iii) separating the solvent through atmospheric distillation until 40% of its initial volume is recovered;
- (iv) cooling the reaction solution to  $20^{\circ}\text{C}$  under stirring and then adding water at a volume adequate to dissolve a halogen salt obtained;
- (v) filtering and purifying the diglycinate by washing with water;
- (vi) drying the methyl diglycinate obtained; and

B) preparing polyglycolyl urea resin:

- (i) stirring together a suspension of cresylic acid and said methyl diglycinate in a reactor at room temperature, stirring until a solution is formed;
- (ii) adding methylene diisocyanate under constant stirring to said solution of said cresylic acid and methyl diglycinate, and keeping temperature of said solution from rising above 60 °C;
- (iii) adding a catalyzer to said solution of ii);
- (iv) raising the temperature of the solution up to 200° C.;
- (v) distilling and then cooling the reaction product; and
- (vi) recovering the polyglycolyl urea resin.

38. (New) The process according to claim 37 wherein the mixture reflux is conducted for 19 hours

39. (New) The process according to claim 37 wherein the resin obtained is cooled to a temperature of 70°C

40. (New) The process according to claim 37, wherein the catalyst in step B(iii) is 1,4 diazobicyclo (2,2,2) octane.

41.(New.) The process according to claim 37 wherein the polyglycolyl urea resin obtained has viscosity (Cp) of 4,800 at 15% solids at 70°C..

42. (New) The process according to claim 37, wherein the  $C_1-C_4$  aliphatic is methanol.

43. (New) The process according to claim 37, wherein the aromatic diglycinate is a methyl diglycinate that corresponds to a stereoisomer mixture having a melting point of  $95-116^{\circ}\text{C}$ .